AE 405 LAB-4

DATA SHEET

GROUP MEMBERS (name, surname, signature)

1.	5.
2.	6.
3.	7.
4.	8.

Procedure

PART (I)

- 1. Fill a glass beaker with (30 ml) of water. Heat the water to 50 C°.
- 2. Determine the temperature of the empty and dry calorimeter before purring the hot water quickly into the calorimeter, stirring continuously for a little while.
- 3. Determine the temperature of the mixture.
- 4. Calculate the capacity of the calorimeter (4).
- 5. Empty the calorimeter, then fill it with water at room temperature, wait a few minutes, then empty and dry the calorimeter.
- 6. Repeat steps (1-4) several times then calculate the average capacity for more accuracy.

PART(II)

- 1. Fill a glass beaker with water.
- 2. Put a known mass piece of metal, tight with a light string 4 cm high from the bottom of the beaker.
- 3. Heat the beaker, water and the piece of metal to T_h = 100 C° for 10 minutes.
- 4. Fill the calorimeter with (ml) of water at room temperature, a few minutes later determines the temperature T_c .
- 5. Drop the hot piece of metal quickly into the calorimeter, stir continuously for a little while then determine the temperature of the mixture T_f.
- 6. Calculate the specific heat of the piece of metal (5).
- 7. Repeat steps (1-6) for the other available pieces of metals.
- 8. Fill in the tables given below.

RESULTS

DATA TABLE-1:

$m_w(gm)$	$c_w \; (cal/g \; c^{\circ})$	$T_c(c^\circ)$	$T_h(c^\circ)$	$T_f(c^\circ)$	$C_{cal} \; (cal/c^{\circ})$

DATA TABLE-2:

Metal	$m_m(gm)$	$m_w(gm)$	$c_w (cal/g \ c^\circ)$	$T_c(c^{\circ})$	$T_h(c^\circ)$	$T_f(c^\circ)$	$c_m (cal/g c^{\circ})$